**Standard Electrode Potentials – Data Page**

Values relative to the standard hydrogen electrode (SHE). (Source: Wikipedia)

* *T*0 = 298.15 K (25.00°C)
* *p*0 = 1.01325 bar = 1 atm (most literature data are at 1 atm, despite the current standard of 1 bar)
* activity *a* = 1 or pure solid, pure liquid, or for water
* for ionic species dissolved in water, *a* = 1 = ** ∙ *c* / *c*0, where *c*0 = 1 mol/L

|  |  |
| --- | --- |
| **Half-cell reaction** | ***E*°  (V vs. SHE)** |
| F2 + 2 e– ⭢ 2 F– | +2.87 |
| Ce4+ + e– ⭢ Ce3+ | +1.61 |
| MnO4– + 8 H+ + 5 e– ⭢ Mn2+ + 4 H2O | +1.51 |
| Cl2 + 2e– ⭢ 2 Cl– | +1.36 |
| O2 + 4 H+ + 4 e– ⭢ 2 H2O | +1.23 |
| MnO2 + 4 H+ + 2 e– ⭢ Mn2+ + 2 H2O | +1.23 |
| Br2 + 2e– ⭢ 2 Br– | +1.06 |
| Ag+ + e– ⭢ Ag | +0.80 |
| Fe3+ + e– ⭢ Fe2+ | +0.77 |
| O2 + 2 H+ + 2 e– ⭢ H2O2 | +0.68 |
| I2 + 2 e– ⭢ 2 I– | +0.54 |
| O2 + 2 H2O + 4e–  ⭢ 4 OH– | +0.40 |
| Cu2+ + 2 e–  ⭢ Cu | +0.34 |
| AgCl(s) + e– ⭢ Ag + Cl– | +0.222 |
| **2 H+ + 2 e–  ⭢ H2** | **0.00** |
| Ni2+ + 2 e–  ⭢ Ni | –0.28 |
| Cd2+ + 2 e–  ⭢ Cd | –0.40 |
| Fe2+ + 2 e–  ⭢ Fe | –0.44 |
| Zn2+ + 2 e–  ⭢ Zn | –0.76 |
| 2 H2O + 2 e–  ⭢ H2 + 2 OH– | –0.83 |
| Al3+ + 3 e–  ⭢ Al | –1.66 |
| Na+ + e–  ⭢ Na | –2.71 |
| Li+ + e– → Li | –3.05 |